

## File Type PDF Chemquest 29 Balancing Equations Answers

# Chemquest 29 Balancing Equations Answers

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Now we will balance the sodium atoms by placing a 6 in front of NaCl.  $\text{Ca}_3\text{N}_2 + 6\text{NaCl} \rightarrow 3\text{CaCl}_2 + 2\text{Na}_3\text{N}$ . Finally, examine the chlorine atoms and notice that they are already balanced. 5. Double check each atom to make sure there are equal numbers of each on both sides of the equation. When balancing equations you NEVER change subscripts.

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Introduction Questions Draw the Lewis  
structures for the following molecules.  
Part c is done for you. a) H<sub>2</sub>O (oxygen is  
central) b) SO<sub>2</sub> c) N<sub>2</sub> It t-t (a [Jrt @ N=N  
@ @ f\ @ 2 In question 1 above you  
should see three different kinds of bonds  
of bonds as best you can. a) single bond:  
one es.iv (ee-) shave a b) \*-l; double  
bond:

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the following molecules.

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Hour I ntroduction Questions  
Information: Subscripts and Coefficients  
A subscript is a small number that tells  
you how many atoms are in a  
compound. For example, in  $\text{CaCl}_2$  the  
two is the subscript and it tells us that  
there are two chloride ions bonded to  
one calcium. A coefficient tells also tells

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us how many atoms or compounds there are, but in a different way.

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exaggeration ways to get this ebook  
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equations is additionally useful. You  
have remained in right site to begin  
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You could ...

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7. Here are the answers to question 6.  
Double check your answers to question  
6: A)  $\text{CaCl}_2 + \text{Li}_2\text{O} \rightarrow \text{CaO} + 2 \text{LiCl}$  B)  
 $\text{Al}_2\text{S}_3 + 3 \text{Cu} \rightarrow 3 \text{CuS} + 2 \text{Al}$  C)  $2 \text{Na}_3\text{P} +$   
 $3 \text{MgBr}_2 \rightarrow \text{Mg}_3\text{P}_2 + 6 \text{NaBr}$  Information:  
Balancing Equations Containing  
Polyatomic Ions When an equation  
contains polyatomic ions, it may look a  
little more difficult to balance. But



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## **Introduction Questions Information: Subscripts and ...**

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neutralization reaction makes water and  
salt hcl naoh h<sub>2</sub> o nacl filename general  
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Balancing Equations: Answers to  
Practice Problems 1. Balanced  
equations. (Coefficients equal to one (1)  
do not need to be shown in your  
answers).

## **Balancing Equations: Practice Problems**

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1 and four times your answer to question 2.  $4 + 4(7) = 4 + 28 = 32$  4. Verify that there are 32 electrons pictured in the drawing of  $\text{CCl}_4$ .

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